

3-(1,2-Diphenylvinyl)-2-methyl-1-phenyl-
sulfonyl-1*H*-indoleK. Palani,^a P. Jaisankar,^b
P. C. Srinivasan^b and
M. N. Ponnuswamy^{a*}^aDepartment of Crystallography and Biophysics,
University of Madras, Guindy Campus, Chennai
600 025, India, and ^bDepartment of Organic
Chemistry, University of Madras, Guindy
Campus, Chennai 600 025, IndiaCorrespondence e-mail:
mnp52004@yahoo.com

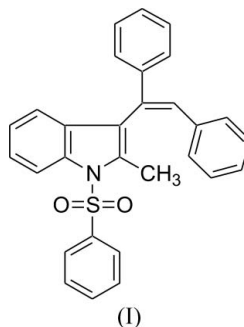
Key indicators

Single-crystal X-ray study
T = 293 K
Mean $\sigma(\text{C}-\text{C}) = 0.006 \text{ \AA}$
R factor = 0.052
wR factor = 0.137
Data-to-parameter ratio = 9.7For details of how these key indicators were
automatically derived from the article, see
<http://journals.iucr.org/e>.

In the title compound, $\text{C}_{29}\text{H}_{23}\text{NO}_2\text{S}$, the bond angles around the S atom indicate a distorted tetrahedral configuration. The crystal structure is stabilized by weak $\text{C}-\text{H}\cdots\text{O}$ and $\text{C}-\text{H}\cdots\pi$ hydrogen bonds.

Comment

The indole unit is observed in many plants (Nigović *et al.*, 2000). Indole-3-acetic acid is a naturally occurring plant growth hormone that controls a number of plant-growth activities (Fargasova, 1994). Many indole-containing natural products are found to exhibit psychotropic (Grinev *et al.*, 1978) and hypertensive (Merk, 1971) properties. A large number of biologically active compounds, mostly those affecting the central nervous system (Zhang & Liebeskind, 1996), contain indolines and their oxidized counterparts as important pharmacophores. Some of the indole derivatives possess anti-tumour (Schollmeyer *et al.*, 1995) and antibacterial (Okabe & Adachi, 1998) activities. In view of this importance, the crystal structure of the title compound, (I), has been determined and the results are presented here.



A *ZORTEP* (Zsolnai, 1997) plot of the molecule is shown in Fig. 1. The S—O, S—C and S—N bond distances are comparable with the previously reported values of 1.435 (5), 1.767 (7) and 1.685 (5) Å, respectively (Govindasamy *et al.*, 1998). As observed in other phenylsulfonylindoles (Rodriguez *et al.*, 1995; Govindasamy *et al.*, 1997, 1998), the C—N distances in the indole ring system are longer due to the electron-withdrawing character of the phenylsulfonyl group. The widening of the O1—S1—O2 angle to 119.4 (2)°, and the concomitant narrowing of the N1—S1—C10 angle to 104.5 (1)°, from the ideal tetrahedral value, are attributed to the Thorpe–Ingold effect (Bassindale, 1984). The C10—C15, C18—C23 and C25—C30 phenyl rings are oriented at angles of 75.1 (1), 81.3 (1), 59.9 (1)°, respectively, with respect to the indole ring system.

C—H \cdots O-type hydrogen bonds are observed in the molecular structure. The crystal structure is stabilized by intermolecular C—H \cdots O and C—H \cdots π hydrogen bonds (Table 2 and Fig. 2).

Received 23 November 2005

Accepted 12 December 2005

Online 21 December 2005

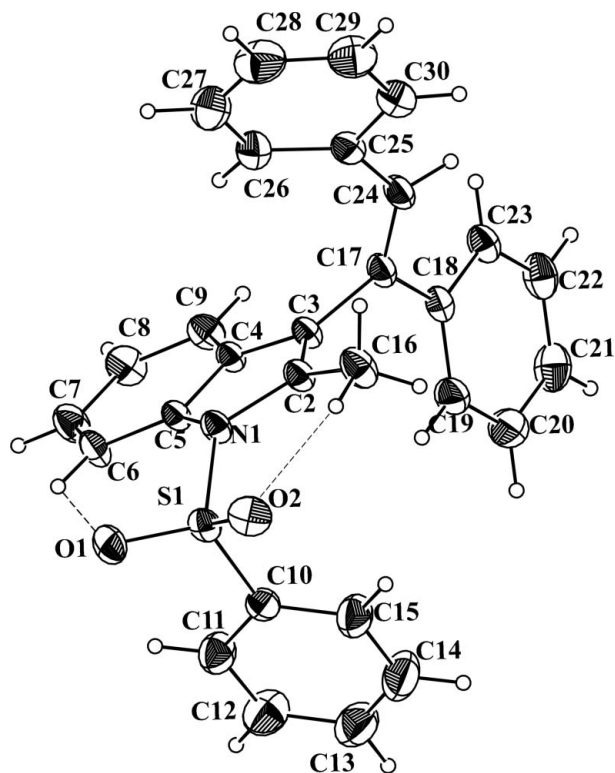


Figure 1
The structure of (I), showing 30% probability displacement ellipsoids and the atom-labelling scheme. Dashed lines indicate C—H...O hydrogen bonds.

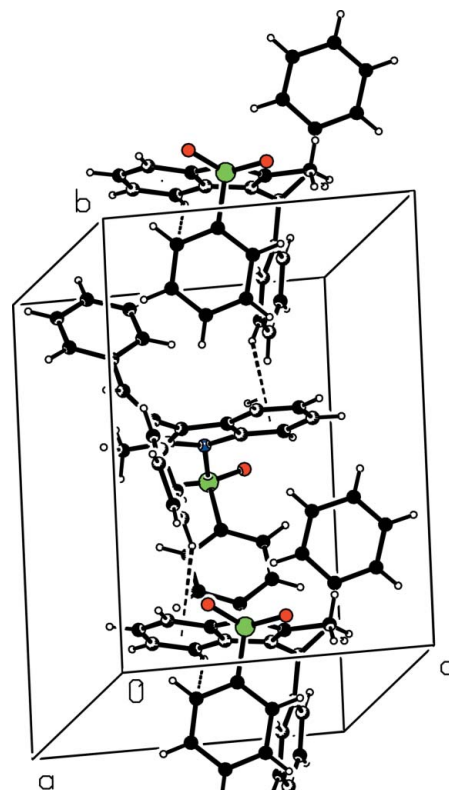


Figure 2
The crystal packing of (I), viewed approximately down the *a* axis. Dashed lines indicate C—H... π interactions.

Experimental

A solution of 3-(1,2-diarylvinyl)-2-methyl-1*H*-indole (5 mmol) in dry tetrahydrofuran (THF, 10 ml) was added slowly to a stirred suspension of 50% sodium hydride (0.24 g, 10 mmol) in dry THF (4 ml) under a nitrogen atmosphere at room temperature. The reaction mixture was refluxed for 3 h and cooled to 268 K. A solution of phenylsulfonyl chloride (1.15 ml, 8 mmol) in dry THF (10 ml) was then added slowly. The solution was then treated with saturated aqueous ammonium chloride solution (50 ml) and the organic layer was separated. The aqueous layer was extracted with chloroform (4 \times 15 ml), and the combined organic layers were dried over anhydrous Na₂SO₄ and concentrated under reduced pressure to give a white oil. This was crystallized from ethyl acetate and hexane (1:4) to give a crystalline solid (yield 1.59 g (71%), m.p. 427 K). ¹H NMR (400 MHz, CDCl₃): *d* = 2.24 (*s*, 3H, CH₃), 6.93–7.58 (*m*, 18H, Ar—H and 1H vinylic), 8.23–8.25 (*d*, 1H, *J* = 8.0 Hz, indole-7H). Mass (*m/z* %): 449 (*M*⁺, 75.6), 308 (77.2), 293 (39.1), 230 (100), 217 (10.8), 77 (34.2), 57 (10.3).

Crystal data

| | |
|---|---|
| C ₂₉ H ₂₃ NO ₂ S | $D_x = 1.272 \text{ Mg m}^{-3}$ |
| $M_r = 449.54$ | Mo $K\alpha$ radiation |
| Monoclinic, $P2_1$ | Cell parameters from 25 reflections |
| $a = 9.029 (3) \text{ \AA}$ | $\theta = 2.3\text{--}28.0^\circ$ |
| $b = 14.891 (7) \text{ \AA}$ | $\mu = 0.16 \text{ mm}^{-1}$ |
| $c = 9.8459 (17) \text{ \AA}$ | $T = 293 (2) \text{ K}$ |
| $\beta = 117.588 (18)^\circ$ | Block, yellow |
| $V = 1173.3 (7) \text{ \AA}^3$ | $0.36 \times 0.28 \times 0.18 \text{ mm}$ |
| $Z = 2$ | |

Data collection

| | |
|--|------------------------------------|
| Enraf–Nonius CAD-4 diffractometer | $\theta_{\text{max}} = 28.0^\circ$ |
| ω scans | $h = 0 \rightarrow 11$ |
| Absorption correction: none | $k = 0 \rightarrow 19$ |
| 3078 measured reflections | $l = -12 \rightarrow 11$ |
| 2905 independent reflections | 3 standard reflections |
| 2520 reflections with $I > 2\sigma(I)$ | frequency: 60 min |
| $R_{\text{int}} = 0.036$ | intensity decay: none |

Refinement

| | |
|---------------------------------|--|
| Refinement on F^2 | $w = 1/[\sigma^2(F_o^2) + (0.1042P)^2 + 0.0294P]$ |
| $R[F^2 > 2\sigma(F^2)] = 0.052$ | where $P = (F_o^2 + 2F_c^2)/3$ |
| $wR(F^2) = 0.137$ | $(\Delta/\sigma)_{\text{max}} = 0.001$ |
| $S = 1.06$ | $\Delta\rho_{\text{max}} = 0.57 \text{ e \AA}^{-3}$ |
| 2905 reflections | $\Delta\rho_{\text{min}} = -0.43 \text{ e \AA}^{-3}$ |
| 299 parameters | Absolute structure: Flack (1983), 0 |
| H-atom parameters constrained | Friedel pairs |
| | Flack parameter: 0.37 (10) |

Table 1

Selected geometric parameters (\AA , $^\circ$).

| | | | |
|-------------|------------|--------------|-----------|
| S1—O1 | 1.423 (2) | S1—C10 | 1.759 (4) |
| S1—O2 | 1.423 (2) | N1—C5 | 1.417 (3) |
| S1—N1 | 1.667 (2) | N1—C2 | 1.436 (3) |
| O1—S1—N1—C5 | −50.4 (3) | C10—S1—N1—C5 | 65.3 (3) |
| O2—S1—N1—C5 | −178.8 (2) | O1—S1—N1—C2 | 168.7 (3) |

Table 2
Hydrogen-bond geometry (Å, °).

| $D-H\cdots A$ | $D-H$ | $H\cdots A$ | $D\cdots A$ | $D-H\cdots A$ |
|--|-------|-------------|-------------|---------------|
| C6—H6···O1 | 0.93 | 2.37 | 2.948 (5) | 120 |
| C9—H9···O2 ⁱ | 0.93 | 2.49 | 3.344 (5) | 153 |
| C16—H16B···O2 | 0.96 | 2.40 | 2.858 (5) | 109 |
| C20—H20···C ₈ ⁱⁱ | 0.93 | 2.83 | 3.658 (5) | 148 |

Symmetry codes: (i) $x+1, y, z$; (ii) $-x+1, y-\frac{1}{2}, -z+1$. C_g denotes the centroid of the C4–C9 benzene ring.

H atoms were placed in idealized positions and allowed to ride on their parent atoms, with C–H = 0.93 or 0.96 Å and $U_{\text{iso}}(\text{H}) = 1.2\text{--}1.5U_{\text{eq}}(\text{C})$.

Data collection: *CAD-4 EXPRESS* (Enraf–Nonius, 1994); cell refinement: *CAD-4 EXPRESS*; data reduction: *XCAD4* (Harms & Wocadlo, 1995); program(s) used to solve structure: *SHELXS97* (Sheldrick, 1997); program(s) used to refine structure: *SHELXL97* (Sheldrick, 1997); molecular graphics: *ZORTEP* (Zsolnai, 1997); software used to prepare material for publication: *PLATON* (Spek, 2003).

KP thanks the University Grants Commission (UGC) Herbal Science programme for financial support under the ‘University with Potential for Excellence’ scheme. The UGC and the Department of Science & Technology (DST) are

gratefully acknowledged for financial support to the Department of Crystallography and Biophysics under the UGC–SAP and DST–FIST programmes.

References

- Bassindale, A. (1984). *The Third Dimension in Organic Chemistry*, ch. 1, p. 11. New York: John Wiley and Sons.
- Enraf–Nonius (1994). *CAD-4 EXPRESS*. Enraf–Nonius, Delft, The Netherlands.
- Fargasova, A. (1994). *Bull. Environ. Contam. Toxicol.* **52**, 706–711.
- Flack, H. D. (1983). *Acta Cryst.* **A39**, 876–881.
- Govindasamy, L., Velmurugan, D., Ravikumar, K. & Mohanakrishnan, A. K. (1997). *Acta Cryst.* **C53**, 929–931.
- Govindasamy, L., Velmurugan, D., Ravikumar, K. & Mohanakrishnan, A. K. (1998). *Acta Cryst.* **C54**, 635–637.
- Grinev, A., Trofimkin, Yu. I., Lomanova, E. V., Andreeva, N. I. & Mashkovskii, M. D. (1978). *Khim. Farm. Zh.* **12**, 80–84. (In Russian.)
- Harms, K. & Wocadlo, S. (1995). *XCAD4*. University of Marburg, Germany.
- Merk, P. (1971). *J. Appl. Phys.* **21**, 62–73.
- Nigović, B., Antolic, S., Kojic-Prodic, B., Kiralj, R., Magnus, V. & Salopek Sondi, B. (2000). *Acta Cryst.* **B56**, 94–111.
- Okabe, N. & Adachi, Y. (1998). *Acta Cryst.* **C54**, 386–387.
- Rodriguez, J. G., del Valle, C., Calderon, C. E. & Ripoll, M. M. (1995). *J. Chem. Crystallogr.* **25**, 249–257.
- Schollmeyer, D., Fischer, G. & Pindur, U. (1995). *Acta Cryst.* **C51**, 2572–2575.
- Sheldrick, G. M. (1997). *SHELXS97* and *SHELXL97*. University of Göttingen, Germany.
- Spek, A. L. (2003). *J. Appl. Cryst.* **36**, 7–13.
- Zhang, D. & Liebeskind, L. S. (1996). *J. Org. Chem.* **61**, 2594–2595.
- Zsolnai, L. (1997). *ZORTEP*. University of Heidelberg, Germany.